

Limit Test for Chlorides

Experiment 1 - Practical Notes

Name: _____ Roll no.: _____ Date: _____

Aim

To compare the impurity response of a test preparation with the prescribed standard under matched conditions.

Principle

Chloride reacts with silver ions to produce silver chloride turbidity: $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}$. The test and standard are compared after the same reagent additions and development conditions. The prescribed acidic medium helps reduce interference from ions such as carbonate.

Requirements

- Prescribed monograph for the test material, impurity standard and suitable matched comparison vessels.
- Silver nitrate test reagent, specified nitric acid and chloride-free vehicle.
- For the modified test: the material-specific approved pretreatment and recovery check.

Procedure

1. Read the monograph

Identify test mass, dissolution method, chloride-standard amount, reagent strengths and comparison time.

2. Prepare test and standard

Dissolve the test portion and prepare the standard with the stated amount of chloride. Use matched cylinders and the same final volume.

3. Develop turbidity

Add the prescribed acid and silver nitrate in the stated sequence. Mix equally and protect from inappropriate light exposure.

4. Compare

At the defined time compare turbidity under the prescribed viewing conditions; record less than, equal to or greater than the standard.

5. Perform the modified test

Use the prescribed substance-specific pretreatment, then carry out the matched comparison. Record how interference was removed and whether the impurity was retained.

Modified Test: What Changes

- The precipitation principle remains the same; preparation is adapted to the named test material.
- For a carbonate-containing matrix, the prescribed acid treatment may remove carbonate with carbon dioxide evolution before the final comparison. This is a matrix example, not a universal modified recipe.
- For coloured or reactive material, use the monograph's approved decolourisation or other pretreatment and a suitable recovery/control check. Apply the same relevant treatment to the comparison preparation when specified.

Observations and Record

Item	Required record	Observation
Test portion / standard	Mass; impurity amount in standard	_____
Conditions	Reagents, final amount and development time	_____
Comparison	Test versus standard response	_____

Worked Example

The numbers below are illustrative; enter your own observations in the record.

- Illustrative limit calculation: a 2.0 g test portion has a stated teaching limit of 20 µg chloride/g. The comparator contains $2.0 \times 20 = 40$ µg chloride.
- A standard containing 20 µg/mL supplies 40 µg in 2.0 mL. These numbers demonstrate comparator preparation, not a universal IP chloride limit.

Result

The test response was ____ compared with the standard. Under valid matched conditions it ____ the prescribed limit; an invalid blank, standard or preparation requires repeat assessment.

Precautions

- Use chloride-free water and clean glassware; contamination can produce a false failure.
- Match reagent additions, final volume and reading time.
- Do not improvise pretreatment with an oxidiser or reducing agent; it may introduce chloride or alter recovery.
- Use the specific substance's prescribed limit; a teaching calculation does not set the permitted impurity concentration.

Viva Questions and Answers

1. What produces turbidity?

Silver chloride particles formed from chloride and silver ions.

2. Why use matched conditions?

The visual response depends on reagent concentration, time and optical conditions.

3. Why can a modified test be necessary?

Colour, insolubility or chemical interference can obscure the ordinary comparison.

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